CHEMISTRY 12

TABLE OF SPECIFICATIONS FOR THE PROVINCIAL EXAMINATION

CURRICULUM			COGNITIVE LEVEL %			
ORGANIZERS	SUB-ORGANIZERS	MARKS SUB-TOTAL	Knowledge	Understanding and Application	Higher Mental Processes	TOTAL %
Reaction Kinetics	A1, 2 Introduction A3, 4, 5 Collision Theory A6, 7, 8 Reaction Mechanisms and Catalysts	12				12
Dynamic Equilibrium	B1, 2 Introduction B3, 4 Le Châtelier's Principle	8				16
	B5, 6 The Equilibrium Constant	8				
Solubility Equilibria	C1, 2, 3 Concept of Solubility C4, 5 Solubility and Precipitation	8				16
	C6, 7, 8 Quantitative Aspects	8				
Nature of Acids and Bases	D1, 2, 3 Properties and Definitions D4, 5, 6 Strong and Weak Acids and Bases	9				9
Acids and Bases: Quantitative Problem Solving	E1, 2 K_w, pH, pOH E3, 4 K_a and K_b Problem Solving	8				8
Applications of Acid-Base Reactions	F1 Neutralizations of Acids and Bases F2, 3 Indicators F4, 5 Hydrolysis of Salts F6, 7 Buffer Solutions F8 Acid Rain	16				16
Oxidation-Reduction	G1, 2 Introduction G3, 4 Balancing Redox Equations	11				11
Applications of Redox Equations	H1, 2 Electrochemical Cells H3 Corrosion H4, 5 Electrolytic Cells	12				12
	TOTAL %	100	11	78	11	100

The values in this table are approximate and may fluctuate.

Examination configuration: 62% in multiple-choice format
38% in written-response format

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